

HOLIDAY HOME WORKS

WORKSHEET [CH-1,2,3,4]

[IT INCLUDES SHORT TYPE AND LONG TYPE BOTH]

CH-1

1. What are the three states of matter? Define each of them with two examples.
2. Define interconversion of states of matter. What are the two factors responsible for the change of states of matter?
3. What happens to water if (a) It is kept in a deep freezer (b) It is heated
Explain the phenomenon of change of state of water.
4. Define:
 - a) Diffusion.
 - b) Brownian motion.
5. (a) State the law of conservation of mass.
(b) What do you observe when barium chloride solution is mixed with a sodium sulphate solution

CH-2

6. Give four differences between physical and chemical changes.
7. Given reason:
 - (a) Freezing of water to ice and evaporation of water are physical changes.
 - (b) Burning of a candle is both a physical and chemical change.
 - (c) Burning of paper is a chemical change.
 - (d) Cutting of a cloth piece is a physical change, though it cannot be reversed.
8. Explain making chapati from flour is both physical and chemical change.
9. Why heating of zinc carbonate is a combined change.

CH-3

10. Differentiate between (a) Pure and impure substances

(b) Homogenous and heterogenous substances

11. How is sodium chloride different from its constituent elements?

12. Give one example for each of the following types of mixtures

(a) Solid-solid homogenous mixture

(b) Solid-liquid heterogeneous mixture

(c) Miscible liquids

(d) liquid-gas homogenous mixture

13. Why is iron sulphide a compound?

CH-4

15. Mention briefly the salient features of Dalton's atomic theory (five points).

16. What are the two main features of Rutherford's atomic model?

17. What are the observations of the experiment done by Rutherford in order to determine the structure of an atom?

18. Define the following terms:

(a) Atomic number

(b) Mass number

(c) Nucleons

(d) Valence shell

19. Name three fundamental particles of the atom. Give the symbol with charge, on each particle.

Q20. Give difference between mass number and atomic mass through one example from each.

Q21. Name the following elements

- (a) An element carrying half the electrons in third shell than second shell
- (b) An element being inert in nature and having four shells

Q22. Determine the following for element "P"

- (a) Atomicity
- (b) Valency
- (c) Formula of oxide
- (d) Electronic configuration