

HYDROGEN SPECTRUM XII- SCIENCE

SUBJECT: (PHYSICS) CHAPTER NUMBER:12

CHAPTER NAME: ATOMS

CHANGING YOUR TOMORROW

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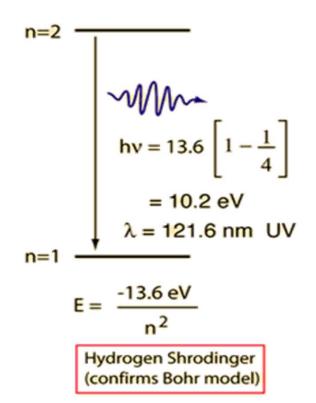
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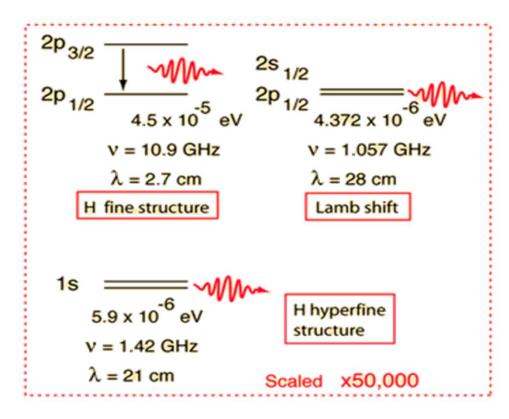
- When an electron jumps amongst energy levels in an atom, energy is emitted or absorbed in the form of electromagnetic radiations and these radiations produce a spectral line of frequencies(or wavelength) associated with an atom, called atomic spectra
- Spectroscopy is the learning and examination emission and absorption spectra associated with an atom to determine its properties
- Spectral lines are the bright and dark line series that constitute the spectrum associated with an atom
- An atom has discrete spectra(where exist a fixed specific lines of energy transition of the electron with discrete energy gaps) also known as quantized spectra
- Another type is continuous spectra(where there are no specific lines of energy transition of electrons) which is the reverse of discrete spectra



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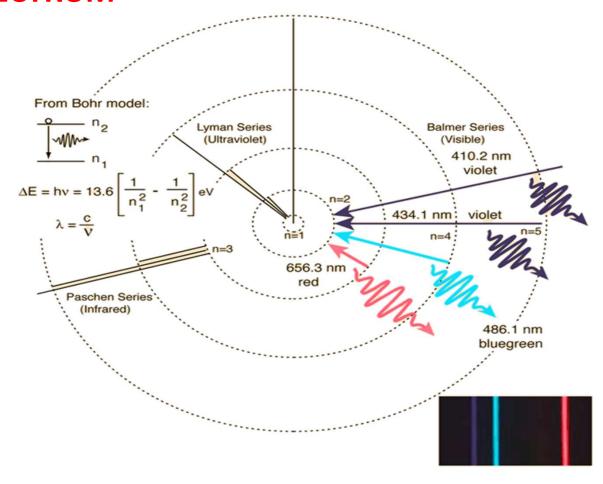
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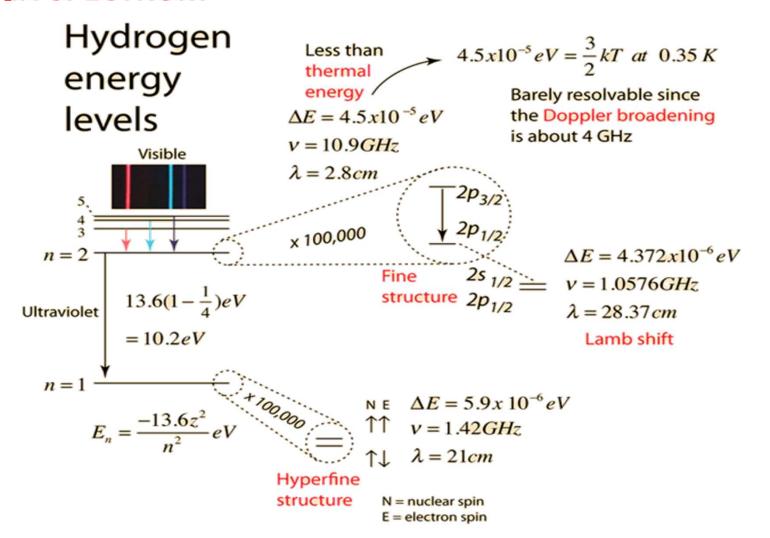
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