

# ATOMS AND MOLECULES

**SUBJECT-CHEMISTRY**  
**CHAPTER NO- 3**  
**Law of Constant Proportion**  
**PERIOD-2**

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**CHANGING YOUR TOMORROW**

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## LEARNING OBJECTIVE

Students will be able to

- Know about the Law of Constant Proportion
- Get aware of the examples proving this law.



# LAW OF CONSTANT PROPORTION

- A pure chemical compound contains the same elements combined together in a fixed proportion by mass is given by the law of definite proportions.
- For e.g., If we take water from a river or from an ocean, both have oxygen and hydrogen in the same proportion.
- It was proposed by the Chemist Joseph Proust in the year 1779.
- For Example, Carbon dioxide can be produced by various ways but the ratio between mass of carbon and oxygen remains fixed always.
- Ratio by mass in  $\text{CO}_2 = 12:32$  OR 3: 8



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# HOME ASSIGNMENT

- Exercise-1 (Q2)
- State and Explain the Law of Constant Proportion.



# THANKING YOU

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