ODM EDUCATIONAL GROUP
Changing your Tomorrow

ODM Teachers' Note

Class	VII		Subject	CHEMISTRY
PD	2	Chapter-3	ELEMENTS, COMPOUNDS AND MIXTURES	
Recapitulation of the previous taught.	✔ We had learnt about the classification of substances or matter. ✔ We had learnt of the elements and its classification. ✔ We had learnt of the Compounds.			
Sub-Concepts	Symbols of Elements. Characteristics of Compounds			
Teaching Aid To be used	Smart Class, PowerPoint presentation, classroom objects, charts.			
Learning Outcome	• Student will be able to know about the symbols of various elements. • They would know about the characteristics of the compounds.			
Sl. No	Step Wise (What to be done)			
1 Introduction	For Achievers Teacher should initiate the discussion on following topics, which will revolve around the core topic of the chapter like, What's your view on the types of substances in this world? ➤ Vision to acquire knowledge of the symbols of the elements. ➤ They need to know of the various rules for writing the symbols. ➤ Vision to acquire knowledge of the characteristics of a compound.		For Average ➤ They would made familiar with the symbols of elements. ➤ They would know of the characteristics of compounds.	

2. Symbols of Elements.		

1 - Hydrogen H	21 - Scandium Sc	41 - Niobium Nb
2 - Helium He	22 - Titanium Ti	42 - Molybdenum Mo
3 - Lithium Li	23 - Vanadium V	43 - Technetium Tc
4 - Beryllium Be	24 - Chromium Cr	44 - Ruthenium Ru
5 - Boron B	25 - Manganese Mn	45 - Rhodium Rh
6 - Carbon C	26 - Iron (Ferrum) Fe	46 - Palladium Pd
7 - Nitrogen N	27 - Cobalt Co	47 - Silver (Argentum) Ag
8 - Oxygen O	28 - Nickel Ni	48 - Cadmium Cd
9 - Fluorine F	29 - Copper (Cuprum) Cu	49 - Indium In
10 - Neon Ne	30 - Zinc Zn	50 - Tin (Stannum) Sn
11 - Sodium (Natrium) Na	31 - Gallium Ga	51 - Antimony (Stibium) Sb
12 - Magnesium Mg	32 - Germanium Ge	52 - Tellurium Te
13 - Aluminium (Aluminum) Al	33 - Arsenic As	53 - Iodine I
14 - Silicon Si	34 - Selenium Se	54 - Xenon Xe
15 - Phosphorus P	35 - Bromine Br	55 - Caesium (Cesium) Cs
16 - Sulfur S	36 - Krypton Kr	56 - Barium Ba
17 - Chlorine Cl	37 - Rubidium Rb	57 - Lanthanum La
18 - Argon Ar	38 - Strontium Sr	58 - Cerium Ce
19 - Potassium (Kalium) K	39 - Yttrium Y	59 - Praseodymium Pr
20 - Calcium Ca	40 - Zirconium Zr	60 - Neodymium Nd

3-Methods for writing the symbols.	Method by which Elements are given Name. ✔ Each element is denoted by a symbol which is usually the first letter of its name in English or Latin. (always in capital letter) ✔ If the first letter of more than two elements is same, the symbol is denoted by two letters, first letter is written in capital while the second letter is written in small letter. ✔ Some symbols have been taken from the Latin, German, or Greek. ✔ These symbols also represent an atom of that element.
4.Characteristics of Compounds.	Elements react to form new compounds. The compound has a fixed composition. Properties of a compound are totally different from those of its constituents. They have a fixed melting point, boiling point, etc. The constituents can be separated only by chemical processes.
5.Home	Exercise-1 Q3(a) ,(b) Q4 & Q5

Assignment	