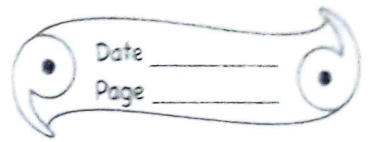


## Ch-3 Worksheet



- 1) The main purpose of separation of mixture is that to remove undesirable substances and to get useful ~~substances~~ and completely pure substances.
- 2) Handpicking is the method ~~decrease~~ of separation of solid solid mixture ~~made~~ in which we remove undesirable substances by hand to get desirable substances.
- 3) The process of separation of grain from husk and hay with help of wind is known as winnowing.
- 4) The method of separation using a magnet separating a mixture of magnetic substance and a non-magnetic substance is called magnetic separation.
- 5) The process in which a solid changes directly into its vapours on heating is called sublimation.
- 6) The solid particle that remains on the filter paper after the filtration. A-Residue

7) The liquid which evaporates and then condenses during the process of distillation - distillate

8) The process of transferring the clean liquid after the solid settles at the bottom of the container.

- Decantation

9) The process by which two miscible liquids are separated - Fractional distillation

10) MCQ

1) A pure liquid is obtained from a solution by:  
Ans. b) Distillation

2) Components of crude petroleum can be separated by:  
Ans. d) Fractional Distillation

3) Examples of a homogeneous mixture is:  
Ans. a) Tap Water

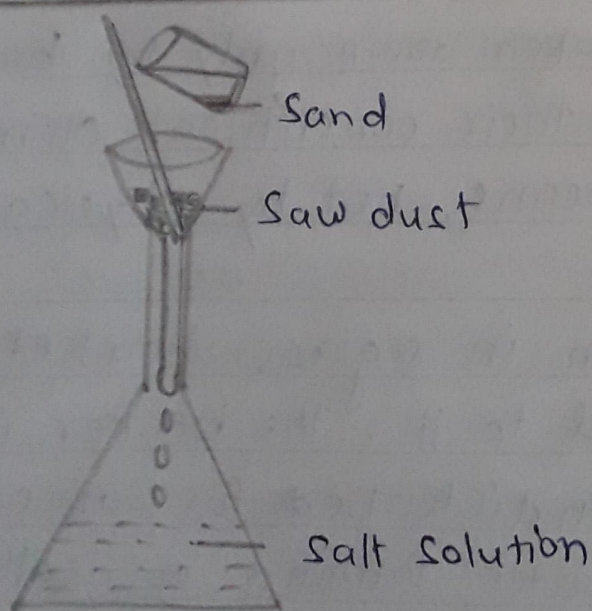
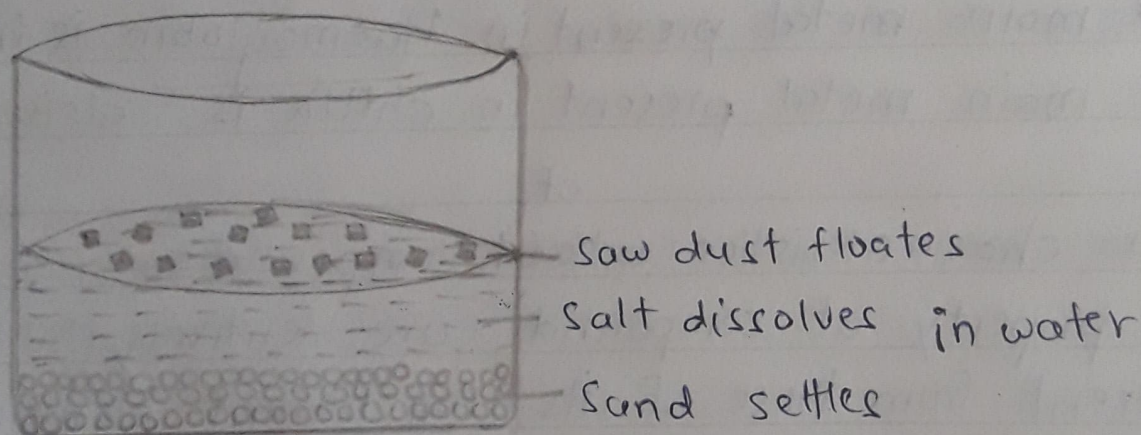
4) In Chromatography the filter paper is:  
Ans. a) Stationery Phase

5) A set of mixtures is:

Ans. a) ink, honey, ice cream, milk

- 1) a) These elements show some properties of metals and some properties of non-metals. They are hard solids. Ex- silicon, Boron, tellurium, polonium etc.
- b) These elements do not react chemically with other elements or compounds, so they are known as noble gases. They are found in air, in traces. They are only six in numbers- helium, neon, argon, krypton, radon and xenon.
- 2) a) The main metal present in haemoglobin is iron.  
b) The main metal present in chalk is calcium.
- 3) Two characteristics ~~without~~ of compound -
- The property of compounds are entirely different from those of its constituent elements.
  - Compounds can be broken down only by broken down ~~only by~~ into their constituent elements only by chemical means, not by physical means.
- 4) The mixture is taken in a glass beaker and water is added to it. The beaker is then allowed to stand. Undisturbed for sometime. Salt dissolves in water, forming a salt solution.

because it is soluble in water. Sawdust being lighter floats on the surface of water while sand being heavier settles down. Now salt solution along with sawdust is decanted slowly on the filter paper fixed in a funnel such that sand is left in a funnel such that sand is left in the ~~beaker~~ beaker as sediment. Salt solution passes through the filter paper while sawdust remains on it. The salt solution is evaporated to get salt from the water. In this way all the components get separated.



5) Crystallisation is a process in which slow evaporation of a solution containing more of the solid component is done. Example - Pure ~~so~~ sugar is obtained from its solution in water by the process of crystallization. At first the sugar solution is heated to evaporate water at a faster speed. When very less of water is left, the solution is cooled. On cooling, the sugar dissolved in it starts separating out in the form of crystals. In evaporation we ~~also~~ does not get the pure form of ~~crystal~~ solid but in crystallization we get pure form of crystal. We usually evaporate the solution until whole of the liquid escapes in its vapour form but in crystallization when very less liquid is present while heating the solution is cooled down.