

Ch-1 MCQs

Q1) One of the following does not happen during a chemical reaction. This is :

Ans) Atoms of one element change into those of another element to form new products because atoms of elements can rearrange but cannot change into other element.

Q2) Which of the following does not involve a chemical reaction?

Ans) Melting of candle wax on heating.

Q3) You are given the solution of lead nitrate. In order to obtain a yellow precipitate, you should mix it with a solution of :

Ans) ~~Potassium~~ Potassium Iodide

Q9) An acid which can decolorize purple colour of potassium permanganate solution is:

Ans) Citric acid

Q5) The chemical reaction between quicklime and water is characterized by:

Ans) Change in temperature

Q6) One of the following is an exothermic reaction.

This is:

Ans) Process of respiration

Q7) The chemical equations are balanced to satisfy one of the following laws in chemical reactions. This law is known as

Ans) Law of conservation of mass

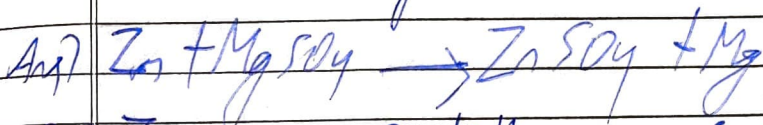
Q8) The chemical reaction between Quicklime and water is characterized by:

Ans) Change in temperature of mixture.



Ans) $4NO_2$

Q10) Which of the following reaction will not take place



Q11) The equation $Cu + xHNO_3 \rightarrow Cu(NO_3)_2 + yNO_2 + zH_2O$, the value of x and y are:

Ans) $x=4$

$y=2$

Q12) When hydrogen sulphide and hydrogen sulphate mix in the presence of water, the reaction $SO_2 + 2H_2S \rightarrow 2H_2O + 3S$ occurs.
Here hydrogen sulphide is acting as:

Ans) A reducing agent

Q13) The layer on iron nails when they are dipped in copper sulphate solution is:

Ans) Hard and shiny

Q14) The substances which combine or react are known as:

Ans) Reactant

Q19) The new substances formed during chemical reactions are known as :-

Ans) Product

Q20) The process of Respiration is :-

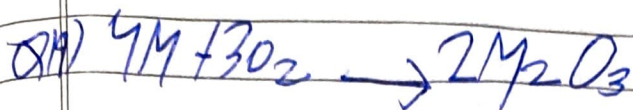
Ans) An oxidation reaction which is exothermic

Q21) In the context of redox reactions, the removal of hydrogen from a substance is known as :-

Ans) Oxidation

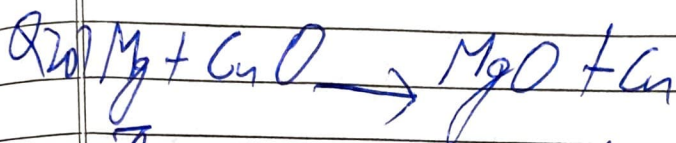
Q22) The chemical reaction involved in the corrosion of iron metal is that of :-

Ans) Oxidation as well as combination



This equation represents:

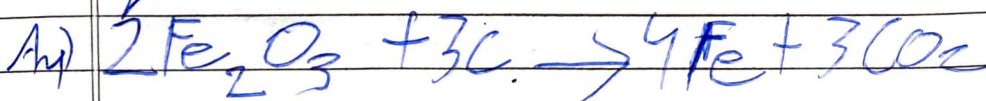
Ans) Combination reaction as well as oxidation reaction



This equation represents

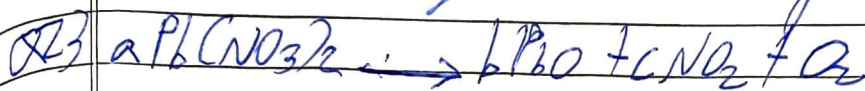
Ans) Redox reaction as well as displacement reaction.

Q21) Which of the following is a balanced chemical equation?



Q1) The term rancidity represents

Ans) ~~oxidation of fats~~ Rottening of food



a, b, c in the above balanced equation is:

Ans) 2, 2, 4

Q3) Corrosion can be prevented by:

Ans) Alloying, Greasing and Galvanizing

Q4) As compare to iron, aluminium has:

Ans) ~~Higher~~ Higher tendency to oxidise.