

## Exercise-III

## OBJECTIVE TYPE QUESTIONS

1. Fill in the blanks:

- (a) Atomicity refers to the number of atoms in the molecule of an element.
- (b) The most abundant element in the earth's crust is oxygen.
- (c) A metal which is a liquid at room temperature is mercury.
- (d) The most abundant element in the atmosphere is Nitrogen.
- (e) A metal which is a poor conductor of electricity is tungsten.
- (f) A diatomic gaseous element is oxygen.
- (g) A liquid non-metal is bromine.

2. Match the columns:

- |     | <u>Column A</u> | <u>Column B</u>                |
|-----|-----------------|--------------------------------|
| (a) | Metals          | (i) Non-reactive               |
| (b) | Molecules       | (ii) Brittle                   |
| (c) | Non-metals      | (iii) Lustrous                 |
| (d) | Noble gases     | (iv) Smallest unit of compound |



3. Indicate whether the following statements are 'true' or 'false'.

(a) A compound is made up of just one kind of atom. **False**

Correct: A compound is made up of two or more elements in a fixed proportion by mass.

(b) Metals reflect light and are good conductors of electricity. **True**

(c) Metals can be polished. **True**

(d) Elements are made up of compounds. **False**

Correct: Elements are made up of atoms.

(e) All elements are artificially prepared.

Correct: All elements are made up of a limited ~~set~~ number of basic substances.

(f) Molecules can exist independently. **True**

(a) Molecules combine to form atoms.

False

Correct Atoms combine to form molecule -

(b) Noble gases are highly reactive. False

Correct Noble gases are non-reactive.

(i) Ozone is a triatomic molecule. True.

### Exercise-II

Q7- Write the molecular formulae of compounds calcium oxide, hydrogen sulphide, carbon monoxide and lead sulphide.

Ans- Calcium oxide is formed of elements calcium (Ca) and Oxygen (O)  
Formulae -  $\text{CaO}$

• Hydrogen sulphide is formed of elements hydrogen (H) and Sulphide (S)  
Formulae -  $\text{H}_2\text{S}$

• Carbon monoxide is formed of elements carbon (C) and Oxygen (O)  
Formulae -  $\text{CO}$



• Lead oxide is made up of elements lead (Pb) and ~~Oxygen (O<sub>2</sub>)~~ Sulphide (S)  
formulae = PbS

Q8 Give two examples each of compounds existing in the following states:

(a) Solid - Table, Chair

(b) Liquid - Water, Juice

(c) Gaseous - Nitrogen, Oxygen

## Extra Question

Q Write formulae of iron oxide, calcium oxide, sodium oxide, zinc chloride

Ans.

Iron oxide -  $\text{FeO}$

Calcium oxide -  $\text{CaO}$

Sodium oxide -  $\text{Na}_2\text{O}$

Zinc chloride -  $\text{ZnCl}_2$