

OBJECTIVE TYPE QUESTIONS.

1. Fill in the blanks:

- a) Atomicity refers to the number of atoms in the molecule of an element.
- b) The most abundant element in the Earth's crust is oxygen.
- c) A metal which is a liquid at room temperature is mercury.
- e) ~~a~~ A metal which is a poor conductor of electricity is tungsten.
- d) The most abundant element in the atmosphere is nitrogen.
- f) A diatomic gaseous element is hydrogen.
- g) A liquid non-metal is bromine.

2. Match Match the columns.

Column A

Column B

- |                |                                   |
|----------------|-----------------------------------|
| a) Metals      | i) Non-reactive (d)               |
| b) Molecules   | ii) Brittle (c)                   |
| c) Non-metals  | iii) Lustrous (a)                 |
| d) Noble gases | iv) Smallest unit of compound (b) |

3. Indicate whether the following statements are true, and false.

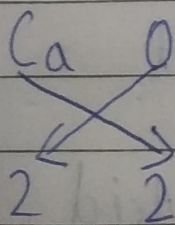
- a) A compound is made up of just one kind of atom. False
- b) Metals reflect light and are good conductors of electricity. True
- c) Metals can be polished. True
- d) Elements are made up of compounds. False

- e) All elements are artificially prepared. False
- f) Molecules can exist independently. True
- g) Molecules combine to form atoms. False
- h) Noble gases are highly reactive. False
- i) Ozone is a triatomic molecule. True

### EXERCISE-II

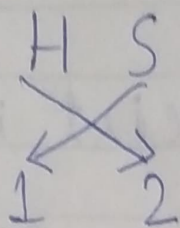
7. Write the molecular formulae of compounds calcium oxide, hydrogen sulphide, carbon monoxide and lead sulphide.

Ans → Compound Calcium oxide is formed of elements calcium and oxygen

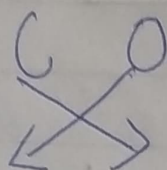


Formula of calcium oxide is  $\text{CaO}$ .

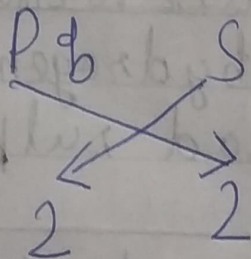
Formula of hydrogen sulphide is  $H_2S$ .



Formula of carbon monoxide is  $CO$ .



Formula of lead sulphide is  $PbS$ .



8) Give two examples each of compounds existing in the following states,

a) Solid - Glucose, Common Salt

b) Liquid - Water, Sulphuric Acid

c) Gaseous - Carbon dioxide, Ammonia

Formulas of

Iron oxide -  $\text{FeO}$

Calcium oxide -  $\text{CaO}$

Sodium oxide -  $\text{Na}_2\text{O}$

Zinc chloride -  $\text{ZnCl}_2$