

EXERCISE-I

1. Homogenous mixtures - Salt solution, Alcohol and water, Air dissolved in water, Air, Sea water, Fruit juices, Brass

Heterogenous mixtures - Petrol, and water, Sand and charcoal, Mist

2. Define the following with an example for each:

a) Pure substance

Ans → Pure substances are either elements or compounds. They contain the same kind of atoms and molecules and have a definite set of physical and chemical properties.

Ex - Gold, Sugar etc.

b) Impure substance

Ans → A substance in which some other substances are also present in smaller or larger amounts is called ~~impure~~ an impure substance.

Ex - Air, Mist etc.

c) Alloy

Ans → A homogenous (solid) mixture of two or more metals or a non-metal is called an alloy.

Ex - Brass, Steel etc.

d) Solution

Ans → The homogenous mixture of water (or any other solvent) and a substance soluble in it (solute) is called a solution.  
i.e. solution = solute + solvent

Ex - Salt solution, Sugar solution etc.

e) Heterogenous mixture

Ans → A mixture in which the components or constituents are not uniformly distributed through its volume and can be easily seen separately is called a heterogenous mixture.

~~f) Homogenous mixture~~  
Ex - Mist, Smoke etc.

f) Homogenous mixture

Ans → A mixture in which its constituents are uniformly distributed through its ~~to~~ volume and ~~to~~ cannot be seen separately is called a ~~homog~~ homogenous mixture etc.

Ex - Milk, Air etc.