Exercise-III
Objective type questions

1. a) Atomicity refers to the number of atoms in the molecule of an element.

b) The most abundant element in the earth's crust is oxygen.

c) A metal which is liquid at room temperature is mercury.

d) The most abundant element in atmosphere is nitrogen.

e) A metal which is a poor conductors of electricity is Tungsten.

f) A diatomic gaseous element is oxygen.

g) A liquid non-metal is bromine.

2.

	Column A	Column B
a)	Metals	iii) Lustrous
b)	Molecules	iv) Smallest unit of compound
c)	Non-metals	ii) Brittle
d)	Noble gases	i) Non-reactive.

3. Indicate whether the following statements are true or false.

a) A compound is made up of just one kind of atom. False.

Correct- ~~Form~~ Element is made up of just one kind of atom. ~~OR~~ or A compound is a substance formed by chemical combination of two or more elements with a fixed proportion by mass.

b) Metals reflect light and are good conductors of electricity. True.

c) Metals can be polished. True

d) Elements are made up of compounds. False

Correct: Elements are made up of atoms.

e) All elements are artificially prepared. False

Correct: All elements are made up of a limited ~~not~~ number of basic substances.

f) Molecules can exist independently. True.

g) Molecules combine to form atoms. False

Correct- Atoms combine to form molecules

h) Noble gases are highly reactive. False

Correct- Noble gases are non-reactive.

i) Ozone is a triatomic molecule, True.

Exercise-II

7. Ans- The molecular formulae of compounds—
calcium oxide — CaO
hydrogen Sulphide — H_2S
carbon monoxide — CO
lead sulphide — PbS

8. Ans- Examples of compounds existing in the states—
* Solids— Sand (Silicon dioxide), Common salt (Sodium Chloride)
* Liquids- Hydrochloric acid and Sulphuric acid and water
* Gaseous— Carbon dioxide and Nitric Oxide

Q. Write formulae of iron oxide, calcium oxide, sodium oxide, zinc chloride.

Ans- iron oxide — FeO
Calcium oxide — CaO
Sodium oxide — NaO
Zinc chloride — $ZnCl_2$

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