

Homework

Q1) Distinguish between the concept of knowing ~~and~~ acids and bases on the basis of Arrhenius Theory and Lewis theory?

Arrhenius theory

* Acids are the hydrogen compound which gives H^+ ions or protons on dissociation in water.

* Bases are the hydroxide compounds which give OH^- ions on dissociation in water.

Lewis theory

* Acid is a substance that accepts a pair of electrons to form a covalent bond

* Base is a substance that donates a pair of electrons to form a covalent bond.

Q2) Although NH_3 doesn't contain any OH^- ions still it behaves as a base why?

NH_3 is a typical weak base. Ammonia itself obviously doesn't contain hydroxide ions but it reacts

with water to produce ammonium ions and hydroxide ions.

Q3) ~~Why~~ What is the ox state of K-atom in potassium permanganate?

Ans The ox-state of K atom is potassium permanganate +1