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1) Arrhenius Acid: Hydronium breaks up up to yields an H^+ in solution.

Arrhenius base: Hydroxide OH^-

is dissolved in water.

Bronsted Lowry acid is a

Acid which donates a Proton

or H^+ ion to the other compound

and forms a conjugated base.

A Bronsted Lowry base is a

Substance which accept a

Proton or H^+ ion to form the

other compound and forms

a conjugate Acid.

2) NH_3 when dissolved in water forms ammonium hydroxide. Ammonium hydroxide ionises to give ammonium ions (NH_4^+) and hydroxide ions (OH^-).

3) The oxidation state of manganese changes as the potassium permanganate decomposes to potassium manganate and manganese dioxide. So, ~~the~~ the oxidation state of K_2MnO_4 and KMnO_4 .

KMnO_4 is +6 and +7 respectively.