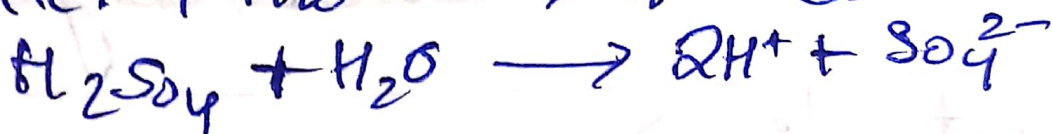


Bases are those chemical substances which react with acids; they form salt & H₂O.

Bronsted & Lowry Theory.
Acids are proton H⁺ donors

Bases are proton H⁺ acceptors



Acid \rightarrow Chemical substance with ~~an~~ which are sour in taste & turn Bl \rightarrow Red

Common Acid \rightarrow vinegar ; HNO_3 ; H_2SO_4 ; HCl

Phy character :- They are sour in taste
 \rightarrow They are watery in touch

\rightarrow Mostly soluble in water at any property.

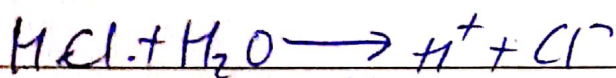
\rightarrow They turn blue litmus to red.

\rightarrow Strong acid are corrosive in Nature
weak (organic Acid) less corrosive.

\rightarrow Their Aq. solutions conducts electricity.

Chemical characteristics \rightarrow

Acids when get dissolved in water, they dissolve in ~~in~~ water they dissociate to give ions which remains. (in the form (H_3O^+))



According to the Arrhenius concept.

Bases: Physical Nature.

- They are bitter in taste
- Slippery in touch; turn Red-Blue
- Their Aq. solution are corrosive in nature
- Dissolve in water.
- Strong Bases are ~~strong~~ weak Acids.

2. Ammonia forms OH^- ions when added to water as it ionised to form ammonium ions & hydroxide ions after forming ammonium hydroxide when mixed in water

$$3. \text{KNO}_3 : 1 + x + (-2) \times 4 = 0$$

$$\Rightarrow 1 + x - 8 = 0$$

$$\Rightarrow x = 7$$