

Bhaswan sarda
Chemistry

Ex-III

1. of Charcoal, soot and coal

b) diamond

c) Copper and aluminium

d) Gold and platinum

e) plastic

2. of machines

b) statues

c) fuel

3. Handle of a frying pan is made up of wood to save our hand as wood is a bad conductor of heat.

b) because it is used for writing

c) because as ~~light~~ bulb needs heat to give light and that's why argon is filled in bulbs

4. a) Because they are ~~bad~~ conductors
of ~~of~~ electrical ~~delivered~~ electricity to
houses, industry, hotel, factory etc and it
is a good conductor of ~~del~~ electricity
b) metals are ~~so~~ soft and material

objective type questions

1. a) compound

b) ~~lit~~ ~~at~~ ~~phere~~ Nitrogen ion

c) 1000°C

d) Nitrogen

e) Tungsten

f) hydrogen

g) bromine

2 a - iii)

b) - iv)

c) - ii)

~~d)~~ d) - 1)

3. ~~a) false~~

b) ~~True~~ False

c) True

d) false.

e) False

f) True

g) ~~True~~ False

h) false

i) ~~True~~

Multiple-choice questions

1. d)

2. e)

3. d)

4. b)

5. b)

d)

7. a)

8. c)

9. b)

10. zinc

b) graphite

c) diamond

d) gold

e) oxygen

Ex-I

1. elements = Mercury, Sulphur, gold, carbon, oxygen, iron, baking soda

compound = Sugar, water, sand, alcohol,

Marble

2. Carbon = C

Calcium = Ca

Copper = Cu

Chlorine = Cl

Cobalt = Co

Argon = Ar

3. A substance of a definite composition which has consistent properties throughout, is called a pure substance. Two types of pure substances are elements and compounds.

• Elements: A substance that has only one molecule called element.
• Compounds: A substance that is made up of only one molecule called

compound.

5. a) metals: mercury and zinc

b) nonmetals: diamond and graphite

c) metalloids: boron and silicon

d) Noble gases: helium and neon

6. Hydrogen and ~~water~~ oxygen.

a) The properties of water are entirely different from those of its constituent elements.

b) water molecule cannot be broken into constituent elements hydrogen and oxygen by simple physical methods.

c) Heat and light are produced when hydrogen and oxygen burn together and chemically combine to ~~form~~ ^{form} water.

7. Metals

1. Metals are lustrous, ductile and malleable.

2. They are good conductors of heat and electricity.

3. They have high melting and boiling points.

Non-metals

1. Non-metals are not lustrous and brittle.

2. They are bad conductors of heat and electricity.

3. They have low melting and boiling points.

8. 1. It consists of two or more elements.

2. The constituting elements of a compound are combined in a definite proportion in every sample of a compound.

3. It is pure and homogeneous.

4. The physical and chemical properties of a compound differ from those of its constituent elements.

9. It is because when two or more elements, when mixed in any proportion and not combining chemically called mixture.

10. a) diamond and sulphur

b) sodium and tin

c) Iodine and carbon

d) Mercury and bromine

e) helium and neon

11. a) sugar :- carbon, hydrogen and oxygen

b) Ammonia :- Nitrogen and hydrogen

c) Marble :- Calcium, carbon and oxygen

d) ~~washing~~ washing soda :- sodium, carbon and oxygen

12 a) H_2O - 1:8

b) CO_2 - 3:8

c) CaO - 5:2

d) $2NO_2$ - 7:16

Ex-11

1. a) A substance that make element and molecules are called atoms.

b) A tiny particle that cannot be seen naked eyes called molecule

e) ~~Atomicity~~ The number of atoms in a molecule of an element is called its atomicity.

d. A formula is a short way of represent the molecules of an element or a compound.

2. Symbols are important to make formula so that formula can mentioned as sub script ~~called~~ to write short form of elements and compounds.

3. Hydrogen = H_2

Oxygen = O_2

Carbon = C

4. A compound like carbon dioxide that has two elements are carbon and oxide.

5. a) 2H represents two atoms of hydrogen while

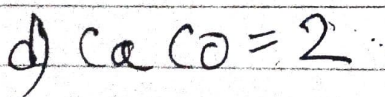
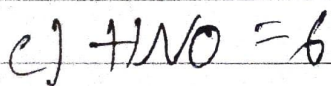
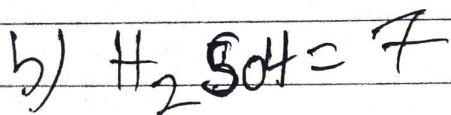
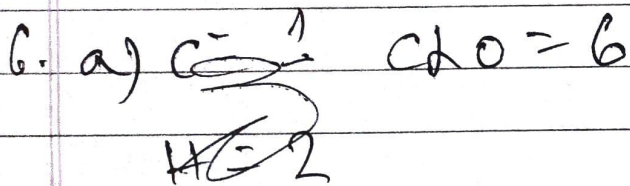
H₂ represents one molecule of hydrogen

b) H₂O represents one molecule of hydrogen

and oxygen while 3 H₂O represent

3 molecules of hydrogen and one

molecule of oxygen.



7. calcium oxide - CaO

Hydrogensulphide - H₂S

Carbon monoxide - CO

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Lead sulphide - PbS

8. a) Solid - Common salt (~~Sodium chloride~~) and
Sand (~~Silicon dioxide~~)

b) Liquid - water and hydrochloric acid

~~c) gas - water and hydrochloric acid~~

c) gas - Carbon dioxide and monoxide