

1) Distinguish between the concept of knowing acids and bases on basis of Arrhenious Theory and Lewis theory

According to Arrhenious Theory acids are the chemical substances that release H^+ ions when dissolved in water.

Bases are the chemical substances that release OH^- ions when dissolved in water.

According to Lewis theory, ~~a~~ bases are the substances that donate electrons
Acids are the substances that gains ~~on~~ receives electrons

2) Although NH_3 doesn't contain any OH^- ions, still it behaves as base. why?

When NH_3 (ammonia) reacts with HCl , NH_4^+ ions and Cl^- ions are formed.



According to Bronsted and Lowery theory, substance which accepts H^+ ions are bases.

Therefore, NH_3 is a base as it accepts H^+ ions.

3) What is the Ox. state of K atom in potassium permanganate?



So, oxidation state of K-atom is +1.