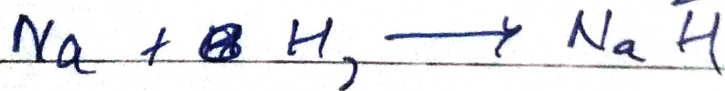
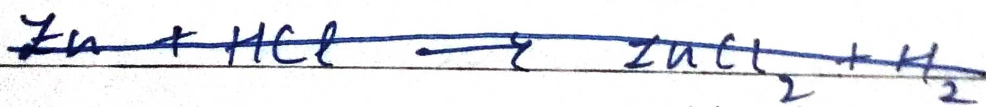
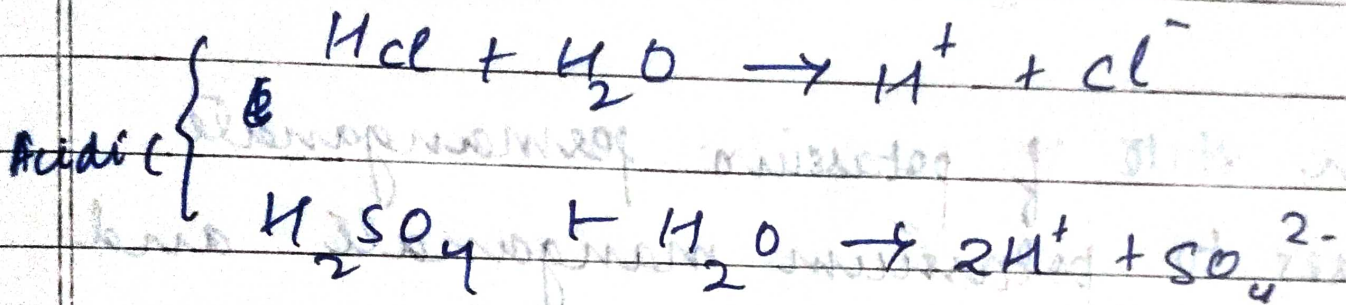
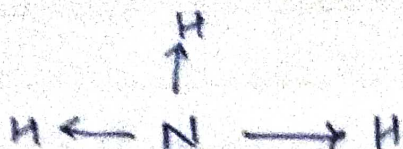


Bronsted & Lowery theory.

Acids are proton's (H⁺) donors

Bases are proton H⁺ acceptors





same protons

Acids are e^- 's acceptors

Bases are e^- 's donors



27 Ammonia doesn't contain any H^+ ions but once it gets dissolved in water it produces H^+ ions. It doesn't fully ~~convert~~ convert the H^+ ions back it is a weak base.

34 Oxidation state of potassium permanganate decomposes to potassium manganate and manganese dioxide.