

H.W
12.08.24

Objective type questions

1. Fill in the blanks.

- a) Atomicity refers to the number of atoms in the molecule of an element.
- b) The most abundant element in the Earth's crust is Oxygen.
- c) A metal which is a liquid at room temperature is mercury.
- d) The most abundant element in the atmosphere is nitrogen.
- e) A metal which is a poor conductor of electricity is tungsten.
- f) A diatomic gaseous element is oxygen.
- g) A liquid non-metal is bromine.

2. Match the columns.

Column A

a) Metals

b) Molecules

c) Non-metals

d) Noble gases

Column B

i) Non-reactive

ii) Brittle

iii) Lustrous

iv) Smallest unit of compound

3. indicate whether the following statements are ~~true~~ true or false.

a) A compound is made up of just one kind of atom.

Ans- False

b) Metals reflect light and are good conductors of electricity.

Ans- True

c) Metals can be polished.

Ans- True

d) Elements are made up of compounds.

Ans- False

e) All elements are artificially prepared.

Ans- False

f) Molecules can exist independently.

Ans- True

g) Molecules combine to form atoms.

Ans- False

h) Noble gases are highly reactive.

Ans- False

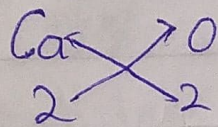
i) Ozone is triatomic molecule.

Ans- True

Exercise II

7. Write the molecular formulae of compounds calcium oxide, hydrogen sulphide, carbon monoxide and lead sulphide.

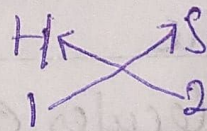
Ans- Compound Calcium oxide is formed of elements calcium (Ca) and oxygen (O).



Symbols combining power. Here subscript number is same Ca_2 Formula of calcium oxide is CaO .

- Compound Hydrogen sulphide is formed of elements hydrogen (H), sulphide (S)

Symbols combining power

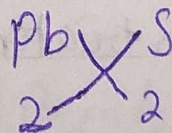


Formula is H_2S .

- Formula of carbon monoxide is CO

- Formula of lead sulphide is PbS .

Symbols combining power -



Here the subscript numbers is same.

8. Give two examples each of compound existing in following states : (a) Solid (b) liquid (c) Gaseous

Ans- (a) Solid - Sand, Common Salt

(b) Liquid - Water, Vinegar

(c) Gaseous - Oxygen, Nitrogen

9. Write formulas of Iron oxide, calcium oxide, Sodium oxide, zinc chloride.

Ans- Formula of Iron Oxide - Fe_2O_3

Formula of Calcium oxide - CaO

Formula of Sodium oxide - Na_2O

Formula of zinc chloride - $ZnCl_2$