

HW
8/7/21

Q1) Explain the terms vaporization and boiling point.

Ans. - The change from liquid state to gaseous state on heating at a constant temperature by absorption of heat is called vaporization or boiling. The temperature at which a liquid changes into vapour without further increase in temperature is called the boiling point of the liquid.

Q2) A liquid can change into vapour state

- At a fixed temperature, and
- at all temperatures

Name the processes involved in the two cases.

Ans. - The processes involved in the two cases are:-

* evaporation

* vaporization

Q3) State three factors which affect the rate of evaporation of a liquid.

Ans. - The factors affecting the rate of evaporation are:-

* The temperature of liquid

* The area of the exposed surface.

* The nature of liquid

Q4) Wet clothes dry more quickly on a warm dry day than on a cold humid day. Explain.

Ans. - In dry air, evaporation is faster than in humid air. This is why wet clothes dry faster.

in dry summer days than in rainy season.
Q5) Why are volatile liquids such as alcohol and spirit stored in tightly closed bottles?

Ans- Volatile liquids with low boiling point such as alcohol, spirit, ether, etc. evaporate much faster than water. This is why volatile liquids are stored in tightly closed bottles.

Q6) Why is cooling produced on evaporation of a liquid?

Ans- The reason for cooling in evaporation is that when a liquid changes into vapour, it requires heat. This heat is supplied by the surroundings of the liquid.

Q7) Explain why in hot summer days water remains cool in earthen pots.

Ans- In summer, water gets cooled in an earthen pot. The reason is that water seeps out on the surface through the pores in the pot and it evaporates. The heat required for evaporation is taken from water inside the pot which therefore gets cooled.