

Metal

Hydrogen
Nitrogen
Chlorine
Phosphorus
Sulphur
Oxygen
Sodium
Iron
Aluminium
Copper

Diatomic
Diatomic
Tetra-atomic
Polyatomic
Diatomic
Monoatomic
Monoatomic
Monoatomic
Monoatomic

Hees

Ex-III

1) Atomicity refers to the number of atoms in the molecule of an element.

- 2) The most abundant element in the earth's crust is mercury
3) A metal which is a liquid at room temperature is mercury
4) The most abundant element in the atmosphere is nitrogen
5) A diatomic gaseous element is oxygen
6) A liquid non-metal is bromine

2) Metals

Non-reactive

b) Molecules

Brittle

c) Non-metals

Lustrous

d) Noble gases

Smallest unit compound

3a) A compound is made up of just one kind of atom.

A = (False)

~~Correct: A compound is made up of two or more~~

b) Metals reflect light and are good conductors of electricity. (True)

c) Metals can be polished. (True)

d) Elements are made up of compounds. (false)

e) All elements are artificially prepared. (false)

f) Molecules can exist independently. (True)

g) Molecules combine to form atoms. (false)

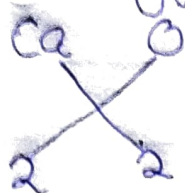
h) Noble gases are highly reactive. (false)

i) Ozone is a triatomic molecule. (True)

Ex II

7 Write the molecular formulae of compounds calcium oxide, hydrogen sulphide, carbon monoxide and lead sulphide.

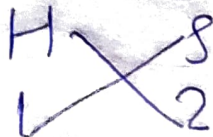
A = Compound calcium oxide is formed of elements calcium (Ca) and oxygen (O)



Symbols combining power Here subscript no. is same as
Formula of calcium oxide is CaO compound hydrogen sulphide is formed of elements Hydrogen (H), sulphide (S)

Symbols combining power

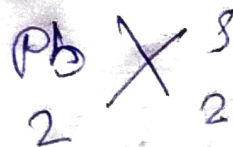
Formula is H_2S



Compound carbon monoxide is compound of elements carbon (C) and oxygen (O) Formula of carbon monoxide is

CO Formula of lead sulphide is PbS

Symbols combining power



Hence the subscript number is same.

8) Give two examples each of compounds existing in the following states:

a) Solid : Table, chair

b) Liquid : Milk, water

c) ~~Solids~~ Gaseous : oxygen, hydrogen

9) Write formulas of iron oxide, calcium oxide, sodium oxide, zinc chloride.

Iron oxide : FeO

Calcium oxide : CaO

Sodium oxide : Na_2O

Zinc chloride : $ZnCl_2$