HW

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① Arhennius Acid

It increases the concentration of H^+ or H_3O^+

Lewis Acid

Acids are electron pair acceptor

Acid is able to form a covalent bond with whatever supplies the electron.

Base (Arhenious)

It increases the concentration of OH^- ions

Lewis Base

Bases are e^- pair donors

② Although NH_3 is a base of not having hydroxyl group it is a base because it forms OH^- ions when added to water as it ionises to form ammonia ions and hydroxide ions after forming ammonium hydroxide when mixed in H_2O

③ $\overset{+1}{K}\ \overset{x}{Mn}\ \overset{-2}{O_4}$

$x = -1 - (-2) \times 4$

$x = -1 + 8 = 7$