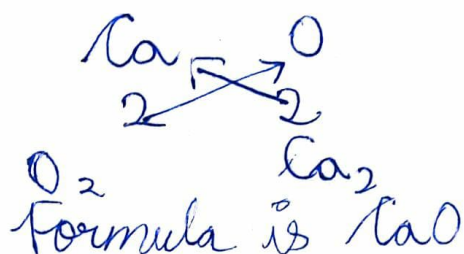
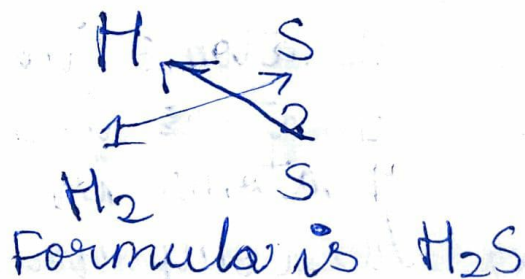


7. Compound calcium oxide is formed of elements calcium (Ca) and oxygen (O).



Hydrogen sulphide is formed of elements hydrogen (H), sulphide (S).



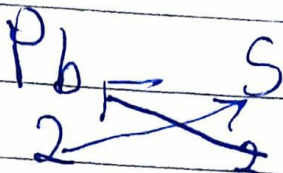
~~For~~

Carbon monoxide is compound of elements carbon (C) and oxygen (O).

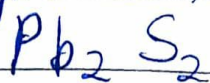
~~Formula of~~



~~Terms~~



Formula of lead sulphide is PbS .
Symbols combining power



8. (a) Solids - Salt, Sugar, Sand
(b) Liquids - Water, Vinegar, Glycerol
(c) Gases - Nitric acid, Sulphuric acid, Hydrochloric acid.

5.9.21

Exercise - III

1. (a) Diamond, Graphite, coal

(b) Diamond is used as gem.

(c) 1. Copper

2. Aluminium as these are shining good conductors of electricity.

(d) 1. Gold.

2. Silver as these are shining, lustrous, and ductile.

(e) Plastic is used as an insulator.

as it is a bad conductor of electricity,

2. (a) Iron: To make machines tools and building material.

(b) Brass: To make water taps and ~~st~~ utensils.

(c) ~~Soil Soil~~

(c) Coal: Coal is used as a fuel also used in thermal power plant to ~~reduce~~ produce electricity.

3. (a) Steel is a good conductor of heat to cook food, pan is made of ~~stelt~~ steel where as wood is insulator of heat and to hold, handle is made up of wood.

(b) Graphite leaves mark on the paper and makes it black.

(c) A is an inert gas and protects the element of bulb from oxidation and burning hence increases bulb's life.

4. (a) Why are copper and aluminium used to make ~~electr~~ electric wires?

Ans:- Copper and aluminium are good conductors of heat and electricity. They can be drawn into wires and beaten into sheets. Therefore, they are used to make electric wires.

(b) Why do you understand by the statement: 'metals are ductile and malleable'?

Ans:- Metals are ductile, i.e., they can be drawn or stretched into thin wires. They are malleable, i.e., they can be beaten into thin sheets.

(c) Give the advantages of using symbols instead of names of elements or compounds.

Ans:- 1. Symbols increase the scientific communication across the world.

2. Symbols help to make equations and data shorter and concise.

3. Symbols are helpful for scientists as it would take time and paper to do their job. They had to write out the full name of every element instead of its symbol.

(d) When iron is mixed with sulphur at room temperature, it does not form a compound.

Why?

Ans:

When iron is mixed with sulphur at room temperature, it does not form a compound ~~because~~ because the mixture of iron and sulphur requires heat to form a compound i.e. ~~the~~ iron sulphide.

(e) Find the atomicity of the following molecules:

1. calcium chloride chloride

2. ~~aluminium sulphide~~ Ammonium sulphide

3. acetic acid

4. dinitrogen oxide

5. ~~carb~~ carbon monoxide

Ans:

~~At~~ Atomicity of

(i) Calcium chloride = ~~Na~~ CaCl_2

$$= (1 \times 1) + (1 \times 2) = 1 + 2 = 3$$

(ii) Ammonium sulphide = $(\text{NH}_4)_2\text{S}$

$$\Rightarrow (1 + 4) \times 2 + 1 \times 1 = 5 \times 2 + 1 \times 1$$

$$\Rightarrow 10 + 1 = 11$$

~~At~~ Atomicity of aluminium sulphide (extra)

Aluminium sulphide = Al_2S_3

$$\Rightarrow (2 \times 1) + (3 \times 1) = 2 + 3 = 5$$

(iii) ~~EA~~
 \Rightarrow

\Rightarrow
(iv) ~~DI~~

\Rightarrow
(v) ~~Car~~

\Rightarrow

1. fill

(a) ~~At~~

(b) oxy

(c) mer

(d) nitro

(e) tung

(f) oxy

(g) per

2. Mat

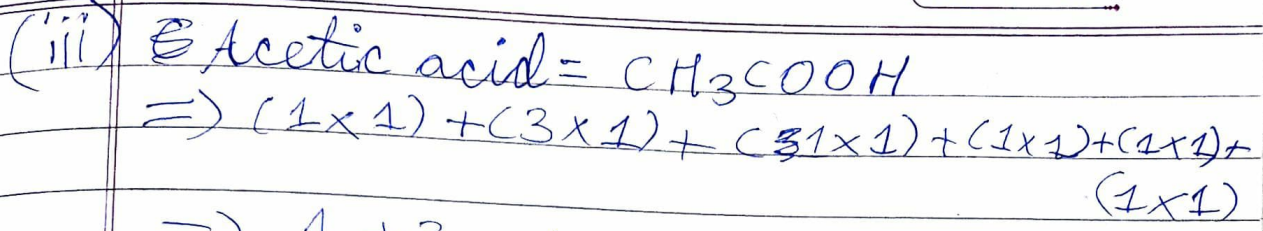
(a) -

(b) -

(c) -

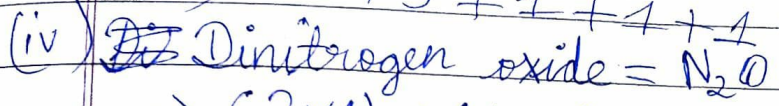
(d) -

room
 sound.



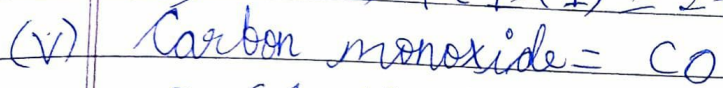
$\Rightarrow 1 + 3 + 1 + 1 + 1 + 1 = 8$

cm



$\Rightarrow (2 \times 1) + (1 \times 1) = 2 + 1 = 3$

s



$\Rightarrow (1 \times 1) + (1 \times 1) = 1 + 1 = 2$

ules:

Objective Type Questions

1. Fill in the blanks: 2. Match:

- | | |
|-----------------------------|-------------|
| (a) At Atomicity | (a) - (iii) |
| (b) oxygen | (b) - (iv) |
| (c) mercury | (c) - (ii) |
| (d) nitrogen | (d) - (i) |
| (e) tungsten | |
| (f) oxygen | |
| (g) bromine | |

3. True / False

(a) False

Correct: A compound is made up of two or more elements in a fixed proportion by mass.

(b) True

(c) True

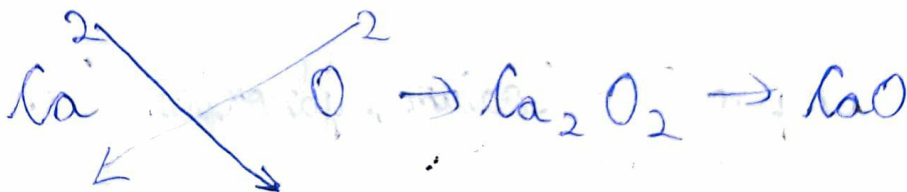
2. Match
- | |
|-------------|
| (a) - (ii) |
| (b) - (i) |
| (c) - (iii) |
| (d) - (iv) |

Homework

Iron Oxide



Calcium Oxide



Sodium Oxide



Zinc Chloride

