

Homework

1. Distinguish between the concept of knowing acids and bases of Arrhenius Theory and Lewis Theory.

Arrhenius Theory

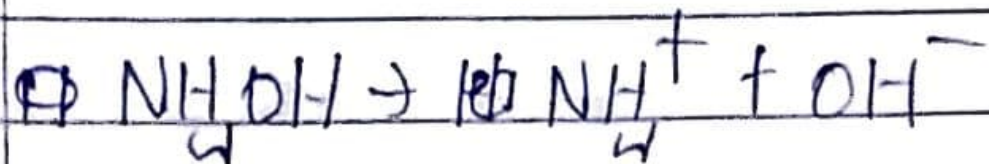
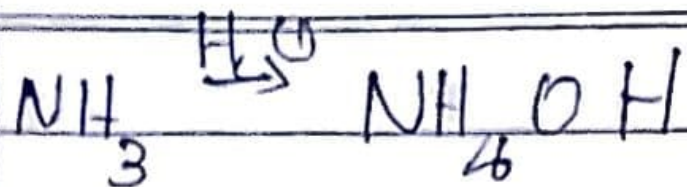
- Acids are proton H^+ donors
- Bases are proton H^+ acceptors

Lewis Theory

- Acids are (e^-) s acceptors
- Bases are (e^-) s donors

2. Although NH_3 doesn't contain any OH^- ions still it behaves as a base. Why?

NH_3 doesn't contain any OH^- ions still it behaves as a base because on reacting with water ammonia forms ammonium hydroxide which further on ionization gives ammonium ion and hydroxide ion.



3. what is the Oxygen state of K-atom in potassium permanganate. 1+