

Homework



1) Distinguish b/w the concept of know acids and bases on basis of Arrhenious theory & Lewis theory

Ans: A/q to Arrhenious theory:

→ An Acid is a substance when dissolved in water, it ionizes and releases hydrogen ions (H^+) in solⁿ.

→ A base is a substance which gives hydroxide/hydroxyl ion (OH^-) in their aq solⁿ.

A/q to Lewis theory.

→ Acids are substances that accept a pair of electron to form a covalent bond.

→ Bases are substances that donate a pair of electrons to form a covalent bond.

2 marks
1) Although NH_3 doesn't contain OH^- ions still it behaves as a base? why?
Ans) Because it forms OH^- ions when added to water and ionizes to form ammonium and hydroxide ions obtaining NH_3OH when mixed with water.

2) What is the oxidation state of K-atom in Potassium Permanganate - $+1$