

26/05

ch-1.

Assignment

Date _____
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Q1. One of the following does not happen during a chemical reaction. This is:

- (a) Breaking of old chemical bonds and formation of new chemical bonds.
- (b) Formation of new substances with entirely different properties.
- (c) Atoms of one element change into those of another element to form new products. (Ans)
- (d) A rearrangement of atoms takes place to form new products.

Q2. Which of the following does not involve a chemical reaction?

- (a) Digestion of food in our body.
- (b) Process of Respiration.
- (c) Burning of candle wax when heated. (Ans)
- (d) Melting of candle wax on heating. (Ans)

Q3. You are given the solution of lead nitrate. In order to obtain a yellow precipitate, you should mix it with a solution of:

- (a) Potassium chloride.
- (b) Potassium Nitride.
- (c) Potassium Sulphate.
- (d) Potassium Iodide. (Ans)

Q4. A acid which can decolorize purple coloured potassium permanganate solution is:

- (a) Sulphuric acid.
- (b) Citric acid. (Ans)
- (c) Carbonic acid.
- (d) Hydrochloric acid.

Q5. The chemical reaction between quicklime & water is characterized by:

- (a) evolution of hydrogen gas.
- (b) formation of slaked lime precipitate.
- (c) Change in temperature of Mixture. (Ans)
- (d) change in colour of product.

Q6. One of following is an exothermic reaction. This is:

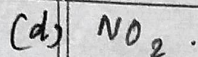
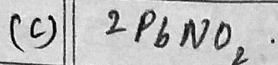
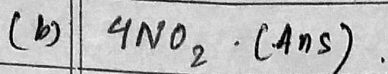
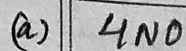
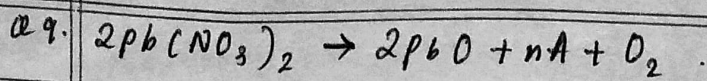
- (a) Electrolysis of water.
- (b) Conversion of Limestone into quicklime.
- (c) Process of Respiration. (Ans)
- (d) Process of photosynthesis.

Q7. The chemical reaction are balanced to satisfy one of following laws in chemical reactions. This law is known as.

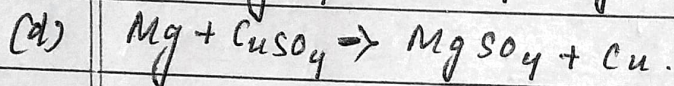
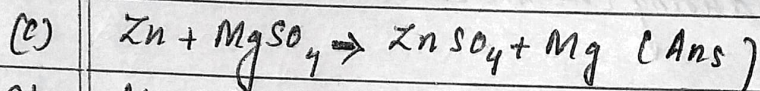
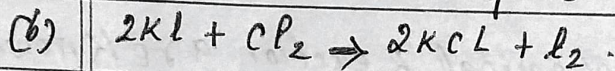
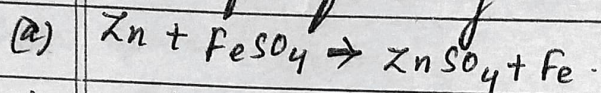
- (a) Law of Conservation of Momentum.
- (b) Law of Conservation of Mass (Ans)
- (c) Law of Conservation of Motion
- (d) Law of Conservation of Magnesium.

Q8. The chemical reaction between Quicklime & water is characterized by:

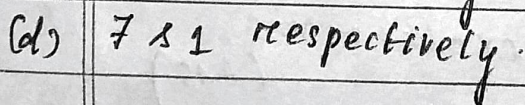
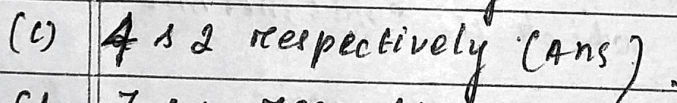
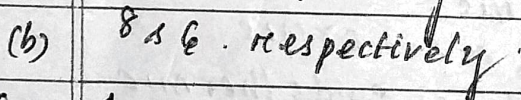
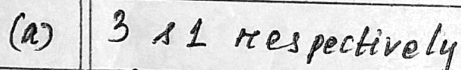
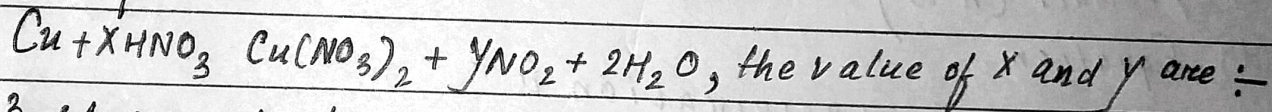
- (a) Evolution of hydrogen gas.
- (b) formation of slaked lime precipitate.
- (c) Change in Temperature of Mixture. (Ans)
- (d) change in colour of product.



Q10. Which of the following reaction will not take place?

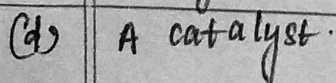
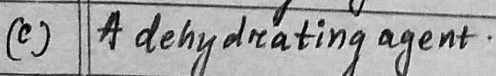
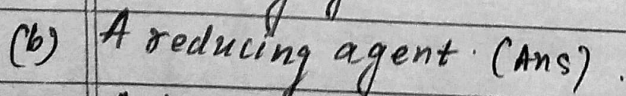
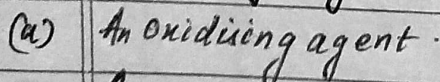


Q11. The equation



Q12. When the gases Sulphur dioxide & Hydrogen Sulphide mix in presence of water, the reaction $\text{SO}_2 + 2\text{H}_2\text{S} \rightarrow 2\text{H}_2\text{O} + 3\text{S}$ occurs.

Here hydrogen Sulphide is acting as:



Q13. The layer on Iron nails when they are dipped in copper sulphate solution is:

- (a) Soft & dull.
- (b) Hard & shiny.
- (c) Hard & dull.
- (d) none of these. (Ans)

Q14. The substances which combine or react are known as:

- (a) Reactant. (Ans)
- (b) production.
- (c) Reaction.
- (d) catalyst.

Q15. The new substances formed during chemical reaction are known as:

- (a) catalyst.
- (b) Reactant.
- (c) Product. (Ans)
- (d) Equation.

Q16. The process of Respiration is:

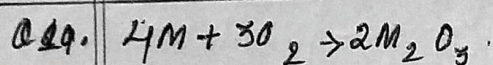
- (a) An Oxidation reaction is endothermic.
- (b) A reduction reaction which is exothermic.
- (c) A combination reaction which is endothermic.
- (d) An Oxidation reaction which is exothermic. (Ans)

Q17. In the context of Redox reactions, the removal of the hydrogen from a substance is known as:

- (a) Oxidation. (Ans)
- (b) Dehydration.
- (c) Reduction.
- (d) Dehydrogenation.

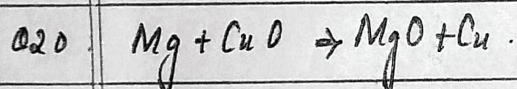
Q18. The chemical reaction involved in the corrosion of iron method is Metal is that of:

- (a) Oxidation as well as displacement.
- (b) Reduction as well as combination.
- (c) Oxidation as well as combination (Ans)
- (d) Reduction as well as displacement.



This equation represents:

- (a) Combination Reaction as well as Reduction reactions.
- (b) Decomposition Reaction as well as Oxidation reactions.
- (c) Oxidation reaction as well as displacement reaction.
- (d) Combination reaction as well as Oxidation reaction. (Ans)



This equation represents:

- (a) Decomposition reaction as well as displacement reaction (Ans)
- (b) Combination reaction as well as double displacement reaction.
- (c) Redox reaction as well as displacement reaction.
- (d) Double displacement reaction as well as Redox reaction.

Q21. Which of the following is a balanced chemical equation?

- (a) $H_2O_2 \rightarrow H_2O + O_2$.
- (b) $2Fe_2O_3 + 3C \rightarrow 4Fe + 3CO_2$.
- (c) $SO_2 + O_2 + 2H_2O \rightarrow 4H_2SO_4$.
- (d) $2Mg + HCl \rightarrow MgCl_2 + H_2$. (Ans).

Q22. The term Rancidity represents.

- (a) Acid Rain
- (b) Oxidation of fatty food ↓
- (c) Rotting of fruit. (Ans)
- (d) fading of the coloured clothes in Sun.

Q23. $a\text{Pb}(\text{NO}_3)_2 \rightarrow b\text{PbO} + c\text{NO}_2 + d\text{O}_2$
 a, b, c in the above balanced equation is:

- (a) 2, 2, 4. (Ans)
- (b) 4, 4, 8.
- (c) 3, 3, 6.
- (d) 2, 2, 2.

Q24. Corrosion can be prevented by:

- (a) Alloying.
- (b) Tinning.
- (c) Galvanizing.
- (d) All of the above. (Ans)

Q25. As compare to Iron, Aluminium has:

- (a) Higher tendency to Oxidise. (Ans)
- (b) lesser tendency to Oxidise.
- (c) Equal tendency to Oxidise.
- (d) none of them.