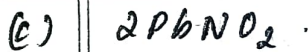
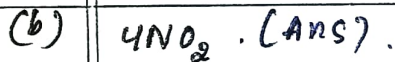
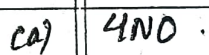
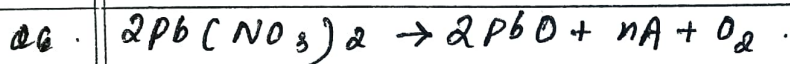


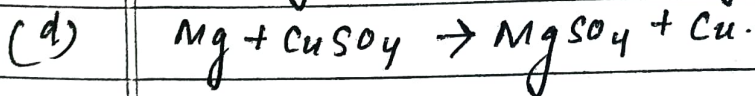
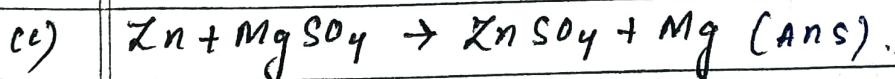
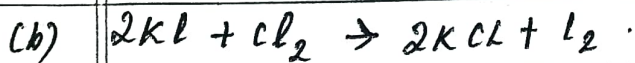
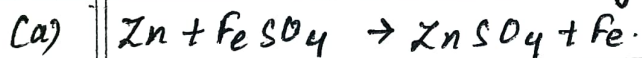
1. Which of the following does not involve a chemical reaction?
- (a) Digestion of food in our body.
  - (b) Process of Respiration.
  - (c) Burning of candle wax when heated.
  - (d) Melting of candle wax on heating (Ans).
- Q2. You are given the solution of lead nitrate. In order to obtain a yellow precipitate, you should mix it with a solution of:
- (a) Potassium chloride.
  - (b) Potassium Nitride.
  - (c) Potassium sulphate.
  - (d) Potassium iodide. (Ans).
- Q3. An acid which can decolorize purple coloured potassium permanganate solution is:
- (a) Sulphuric acid.
  - (b) Citric acid. (Ans).
  - (c) Carbonic acid.
  - (d) Hydrochloric acid.
- Q4. One of the following is an exothermic reaction. This is:
- (a) Electrolysis of water.
  - (b) Conversion of limestone into quicklime.
  - (c) Process of respiration. (Ans).
  - (d) Process of photosynthesis.

Q6. The chemical reaction (Equation) are balanced to satisfy one of the following laws in chemical reactions. This law is known as.

- (a) law of conservation of momentum.
- (b) law of Conservation of mass. (Ans)
- (c) law of Conservation of motion.
- (d) law of Conservation of magnesium.



Q7. Which of the following reaction will not take place?



Q8. The layer on iron nails when they are dipped in Copper Sulphate solution is:

(a) Soft & dull.

(b) Hard & shiny.

(c) Hard & dull. (c) (Ans).

(d) none of these.

Q9. The substances which combine are known as:

- (a) Reactant. (Ans)
- (b) production.
- (c) reaction.
- (d) Catalyst.

Q10. The new substances formed during chemical reaction are known as:

- (a) catalyst.
- (b) reactant.
- (c) product. (Ans)
- (d) reaction.

In the context of Redox Reaction, the removal of hydrogen from a substance is known as:

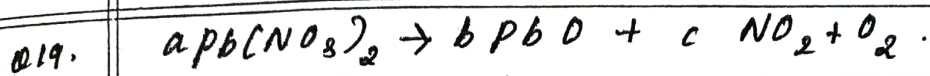
- (a) Oxidation.
- (b) Dehydration.
- (c) Reduction. (Ans)
- (d) Dehydrogenation.

Q12. Which of the following is a balanced chemical equation?

- (a)  $H_2O_2 \rightarrow H_2O + O_2$ .
- (b)  $2Fe_2O_3 + 3C \rightarrow 4Fe + 3CO_2$ .
- (c)  $SO_2 + O_2 + 2H_2O \rightarrow 4H_2SO_4$ .
- (d)  $2Mg + HCl \rightarrow MgCl_2 + H_2$ . (Ans)

Q13. The term Rancidity represents.

- (a) Acid Rain.
- (b) Oxidation of fatty food (Ans)
- (c) Rottening of fruit.
- (d) fading of the coloured clothes in the sun.



a, b, c in the above balanced equation is :

(a) 2, 2, 4. (Ans)

(b) 4, 4, 8.

(c) 3, 3, 6.

(d) 2, 2, 2.

Q20. Corrosion can be prevented by :

(a) Alloying.

(b) Tinning.

(c) Galvanizing.

(d). All of these. (Ans).

Q21. As compare to Iron, Aluminium has :

(a) Higher tendency to oxidise. (Ans).

(b) Lesser tendency to oxidise.

(c) equal tendency to oxidise.

(d). none of these.

—●— End —