1) Distinguish between the concept of knowing acids and bases on basis of Arrehenious theory & Lewis theory.

According to Arrehenious theory,

→ An acid is a substance which dissolved in water, it ionizes and releases hydrogen ions ($H^+_{(aq)}$) in sol^n.

→ A base is a substance which give hydroxide / hydroxyl ion (OH^-) in their aq. sol^n.

According to Lewis theory,

→ Acids are substances that accepts a pair of electron to form a covalent bond.

→ Bases are substances that donate a pair of electrons to form a covalent bond.

c) Although NH_3 doesn't contain any OH^- ions still it behaves as a base. Why?

Because it forms OH^- ions when added to water as it ionises to form ammonium ions and hydroxide ions after forming NH_4OH when mixed with water.

d) What is the Oxidation state of K-atom in pottasium permanganate. +1