

becomes greater than the weight of the balloon. Hence, the balloon rises up under the influence of net upward force.

[G] Numericals:

① (a) $1 \text{ l} = 1000 \text{ cm}^3$
(1000 mL)

$\Rightarrow 1.28 \text{ g} \Rightarrow 1000 \text{ cm}^3$

$\Rightarrow 1 \text{ cm}^3 = \frac{1}{1000} \text{ g}$

$\Rightarrow \frac{1.28}{1000} = \text{g}$

$\Rightarrow 0.00128 \text{ g/cm}^3$

$\Rightarrow 0.00128 \text{ g cm}^{-3}$

(b) $1.28 - 1 \text{ l}$
 $1 \text{ l} = 1000 \text{ cm}^3$

$\left[\begin{aligned} 1 \text{ cm}^3 &= \frac{1}{100} \\ 1 \text{ cm}^3 &= \frac{1}{1000000} \text{ m}^3 \end{aligned} \right.$

$\left. \begin{aligned} 1000 \text{ cm}^3 &= 1 \text{ m}^3 \\ &\quad \uparrow \\ &\quad 1000 \end{aligned} \right]$

$= \frac{1.28 \text{ kg}}{1000} \div \frac{1 \text{ m}^3}{1000}$

$= 1.28 \text{ kg/m}^3$

$$\textcircled{2} \text{ Density} = 1.11 \text{ kg m}^{-3}$$

$$\text{Volume} = (L \times B \times H)$$

$$= 10 \text{ m} \times 7 \text{ m} \times 5 \text{ m}$$

$$= 350 \text{ m}^3$$

$$\text{Mass} = \text{Density} \times \text{Volume}$$

$$= 1.11 \times 350$$

$$= \underline{\underline{388.5 \text{ kg}}}$$

$$\begin{array}{r} 350 \\ \times 1.11 \\ \hline 350 \\ 350 \times \\ + 350 \times \\ \hline 388.50 \end{array}$$

$$\textcircled{3} 1 \text{ g} = \frac{1}{1000} \text{ kg}$$

$$2.7 \text{ g} = \frac{2.7}{1000} = 2.7 \times 10^{-3} \text{ kg}$$

$$1 \text{ cm}^3 = 1000000 \text{ m}^3$$

$$= 10^{-6} \text{ m}^3$$

$$\text{kg m}^{-3} = \frac{2.7 \times 10^{-3} \text{ kg}}{10^{-6} \text{ m}^3}$$

$$\text{Density of aluminium} = 2700 \text{ kg m}^{-3}$$

$$= 2700 \text{ kg m}^{-3}$$

$$\begin{array}{r} 2.7 \times 10000 \\ \hline 1000000 \end{array}$$

4) $1 \text{ kg} = 1000 \text{ g}$
 $600 \text{ kg} = 600 \times 1000$
 $= 600000 \text{ g}$

$1 \text{ m}^{-3} = 100 \text{ cm}^{-3} = 100000 \text{ cm}^{-3}$

$1 \text{ g/cm}^{-3} = \frac{600000}{100000}$
 $= 0.60 \text{ g/cm}^{-3} = 0.60 \text{ g cm}^{-3}$

5) Mass of the zinc = 438.6 g
 Volume = ~~200 cm^3~~ 86 cm^3

Density = $\frac{\text{Mass}}{\text{Vol.}} = \frac{438.6}{86}$
 $= 5.1 \text{ g cm}^{-3}$

$\begin{array}{r} 84 \ 438 \\ 430 \\ \hline 86 \\ 256 \\ \hline 0 \end{array}$

6) (as) Mass = 150 g
 Volume = 200 cm^3

Density = $\frac{\text{Mass}}{\text{Volume}} = \frac{150}{200}$

$= 0.75 \text{ g/cm}^{-3}$
 $= 0.75 \text{ g cm}^{-3}$

$\begin{array}{r} 200 \ 1500 \\ 1400 \\ \hline 1000 \\ 2000 \\ \hline 0 \end{array} \quad 0.75$

$$(e) \text{ g/cm}^3 \rightarrow \text{kg/m}^3$$

$$\Rightarrow \frac{0.75}{1000} \div 1000000$$

$$\Rightarrow \frac{0.75}{1000} \times \frac{1}{1000000}$$

$$= 0.0075 \times \frac{1}{100000}$$

$$= \frac{0.75 \text{ g}}{1000} \downarrow$$

$$\Rightarrow \frac{0.75}{1000} \div \frac{1}{1000000}$$

$$\Rightarrow \frac{0.75}{1000} \times 1000000$$

$$\Rightarrow 0.75 \times 1000$$

$$\Rightarrow 750 \text{ kg m}^{-3}$$

$$1000 \text{ g} = 1 \text{ kg}$$

$$0.75 \text{ g} =$$

$$\frac{0.75}{1000} \text{ kg}$$

$$1 \text{ cm} = \frac{1}{100} \text{ m}$$

$$1 \text{ cm}^2 = \frac{1}{100} \text{ m}^2$$

$$1 \text{ cm}^3 = \left(\frac{1}{100} \times \frac{1}{100} \times \frac{1}{100}\right) \text{ m}^3 = \frac{1}{1000000} \text{ m}^3$$

$$= 0.000001 \text{ m}^3$$