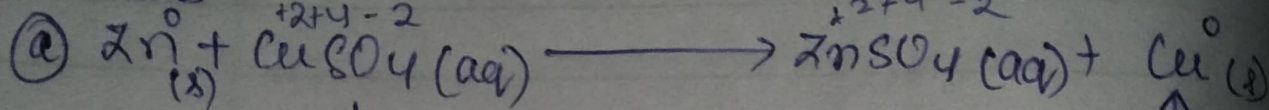


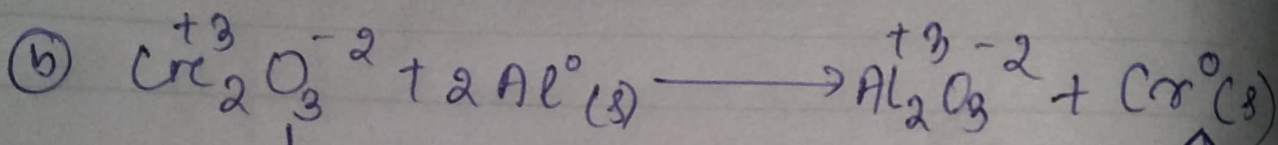
① Give two examples of displacement reactions of both metals and non-metals.

Displacement reaction of:

(i) Metals:

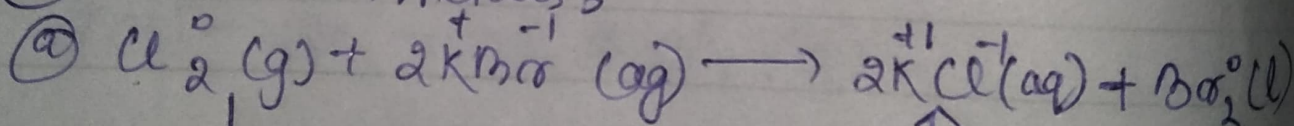


Cu gets reduced from +2 O.S. to 0.

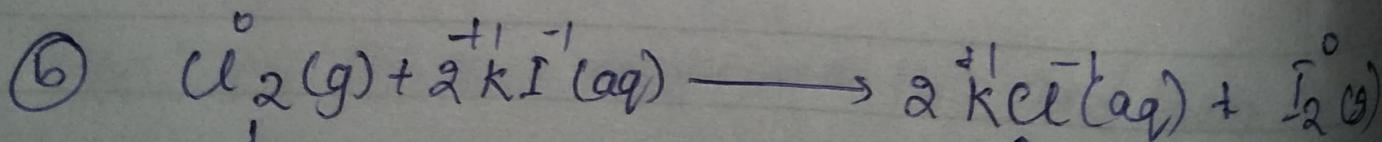


Cr gets reduced from ~~3~~ 3 O.S. to 0.

(ii) Non ~~metals~~ metals:



Cl gets reduced from 0 to -1 O.S.

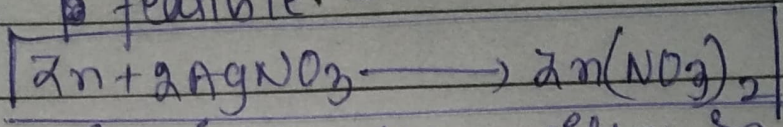


Cl gets reduced from 0 to -1 O.S.



(2) write the feasibility of reactions of the following:

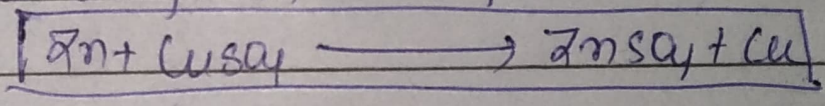
(i) Zinc granules with  $\text{AgNO}_3$ :  
Yes, it is feasible.



Since, zinc is higher than silver in the reactivity series, so it displaces silver. Thus the displacement reaction is feasible.

(ii) Zinc granules with  $\text{CuSO}_4$ :  
Yes it is possible.

Since, zinc is higher than copper in reactivity series, it displaces Cu from  $\text{CuSO}_4$  solution.



(iii) Zinc granules with aluminium sulphate:  
It is not possible.

This is because, aluminium is more reactive than zinc according to the reactivity series. So, zinc cannot displace aluminium.

So, it is not possible.