

# HOMEWORK

## OBJECTIVE TYPE QUESTIONS

Date \_\_\_\_\_

Page \_\_\_\_\_

EXERCISE - III

SURYA NARAYAN MOHANTY

1) Fill in the blanks

- a) Atomicity refers to the number of atoms in the molecule of an element.
- b) The most abundant element in the earth's crust is Oxygen.
- c) A metal which is a liquid at room temperature is Mercury.
- d) The most abundant element in the atmosphere is nitrogen.
- e) A metal which is a poor conductor of electricity is tungsten.
- f) A diatomic gaseous element is Oxygen.
- g) A liquid non-metal is bromine.

2) Match the Columns:

- a) Metals - Lustrous
- b) Molecules - Smallest unit or Compound
- c) Non-metals - Brittle
- d) Noble gases - Non-reactive

Q3) Indicate whether the following statements are true or false.

a) A compound is made up of just one kind of atom. False

b) Metals reflect light and are good conductors of electricity. True

c) Metals can be polished. True

d) Elements are made up of compounds. False

e) All elements are artificially prepared. False

f) Molecules can exist independently. True

g) Molecules combine to form atoms. False

h) Noble gases are highly reactive. False

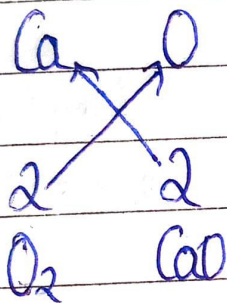
i) Ozone is a triatomic molecule. True

Date \_\_\_\_\_  
Page \_\_\_\_\_

Exercise <sup>(11)</sup> = 7/8

Q7) Write the molecular formulae of Compounds Calcium oxide, hydrogen sulphide, Carbon monoxide and lead sulphide.

Answer  $\Rightarrow$  Compound  $\otimes$  Calcium oxide is formed of element Calcium (Ca) and Oxygen (O).



Symbols Combining power Here the Subscripts number is same  $\text{Ca}_2$  formula of Calcium oxide is  $\text{CaO}$ .

Q8) Give two example each of Compounds existing in the following States:

(a) Solid (b) Liquid (c) Gaseous

a) Solid = Book, Gold

b) liquid = Water, Oil

c) Gaseous = Petrol, Oxygen