EXERCISE III

1. Fill in the blanks:

(a) atomicity refers to the number of atoms in the molecule of an element.

(b) The most abundant element in the earth's crust is Oxygen.

(c) A meatal which is a liquid at room temperature is mercury.

(d) The most abundant element in the atmosphere is nitrogen.

(e) A metal which is a poor conductor of electricity is tungsten.

(f) A diatomic gaseous element is

Oxygen.

(g) A liquid nonmetal is bromine.

2. Match the column.

(a) Metals - lustrous
(b) Molecules - Smallest unit of compound
(c) Non-metals - Brittle
(d) Noble gases - non reactive

3. (a) A compound is made up of just one kind of atom. False

(b) Metals reflect light and are good conductors of electricity. True

(c) Metals can be polished. True

(d) Elements are made up of compounds. False

(e) All elements are artificially prepared. False

(f) Molecules can exist independently. True

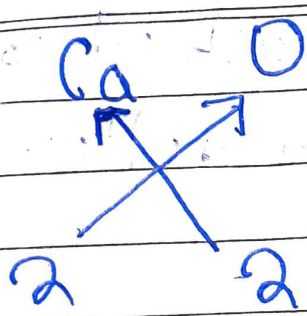
(g) Molecules combine to form atoms. False

(h) Noble gases are highly reactive. False

(i) Ozone is a triatomic molecule. True

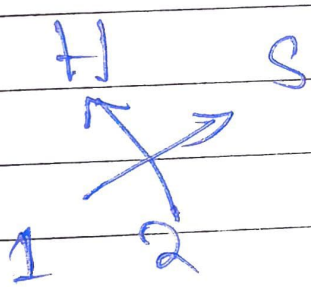
EX= O_3

7. Write the molecular formulae of compounds ~~existing in the following states~~ calcium oxide, hydrogen sulphide, carbon monoxide and lead sulphide.



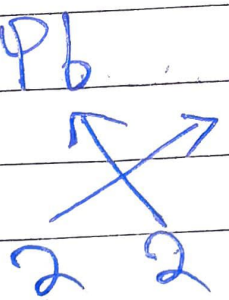
O_2 Ca O

Hydrogen Sulphide



Formula is H_2S

Carbon monoxide = CO



– Lead sulphide

Pb_2S_2 PbS

8. Give two examples each of compounds existing in the following states:

(a) Solid (b) liquid (c) Gaseous

solid - Sodium chloride, Pottassium chloride

liquid - Water, Sodium hydroxide

Gas: Carbon dioxide, Sulfur dioxide

Q. Write formulas of iron oxide, calcium oxide, Sodium oxide, zinc chloride

Ans: Iron oxide - FeO

calcium oxide - CaO

Sodium oxide - Na_2O

zinc chloride - $ZnCl_2$