

HW
18/11/23

ODM CONNECT APP HOMEWORK

1. Give five major differences between elements and compounds.

Ans) Elements

- Cannot be broken down to simpler substances.
- An element represents only one type of atoms eg- sulphur atoms
- An element is made up of one type of atoms.
- Elements have certain characteristic properties of their own.
- Can form diff. types of chemical bonds depending on stability, electron configuration - ions.

Compounds

- Have fixed composition and can be broken down into elements by chemical or electrochemical reactions.
- A compound represents only one type of molecules eg- chlorine molecules.
- A compound may be made of same or different atoms eg- oxygen molecule, carbon dioxide molecule.
- Constituting elements in a compound lose their own characteristics.
- Can have covalent bonds, ionic bonds or metallic bonds.

2. Explain, why we say H_2O is a compound but H_2 and O_2 are elements?

Ans) In compounds, the constituting elements lose their own properties during formation. So that is the reason for which the properties of compound are very different from its constituting elements. The composition of pure water is same throughout. It always contains one part of hydrogen and eight parts of oxygen by weight. Water is formed when an electric spark is passed through the mixture of hydrogen and oxygen. The properties of water thus formed are altogether different from those of its constituent elements. For eg, hydrogen is a combustible gas while oxygen is a supporter of combustion. Water on the other hand is neither combustible nor a supporter of combustion, it actually extinguishes the fire.